

The Crystal Structure of Barium Telluropentathionate Trihydrate

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The crystal structure of barium telluropentathionate trihydrate, $\text{BaTe}(\text{S}_2\text{O}_3)_3 \cdot 3\text{H}_2\text{O}$, has been determined by X-ray methods. The salt crystallizes in the monoclinic space group $P2_1/c$ (No. 14) with four formula units in a unit cell of dimensions $a = 11.139(3)$ Å, $b = 5.243(2)$ Å, $c = 21.306(6)$ Å, $\beta = 106.94(4)^\circ$. The refinement was carried out by a full-matrix least squares program using 2389 non-zero reflections, recorded by means of a single-crystal diffractometer. The R value was 0.024.

The telluropentathionate ion has the *trans* form, the sulphonate groups being rotated out of the plane of the three middle atoms to different sides of the plane. The dimensions of the S-S-Te-S-S chain, from one end of the chain to the other, are $\text{S}(1)-\text{S}(2) = 2.102(2)$ Å, $\text{S}(2)-\text{Te} = 2.384(2)$ Å, $\text{Te}-\text{S}(4) = 2.392(2)$ Å, $\text{S}(4)-\text{S}(5) = 2.090(2)$ Å, $\angle \text{S}(1)-\text{S}(2)-\text{Te} = 103.08(5)^\circ$, $\angle \text{S}(2)-\text{Te}-\text{S}(4) = 96.02(4)^\circ$, $\angle \text{Te}-\text{S}(4)-\text{S}(5) = 102.59(5)^\circ$. The SSte/STeS dihedral angles are 87.7° and 98.7° .

The crystal structures of two barium pentathionate dihydrates,^{1,2} of barium selenopentathionate dihydrate^{3,4} and trihydrate,⁵ and of barium telluropentathionate dihydrate^{6,7} and trihydrate⁸ have been reported. The latter structure was solved and refined by means of film data. The results of the refinement were not quite satisfactory, the standard deviations being high, compared to those of the barium selenopentathionate hydrates.^{4,5} The present work is a more detailed description of this structure analysis, including least squares refinement of diffractometer data.

The pentathionate, selenopentathionate, and telluropentathionate ions have been found to occur in two rotational-isomeric forms, and crystallize in the *cis* form in all of their barium salts, except in barium telluropentathionate trihydrate. The telluropentathionate ion has the *trans* form in the trihydrate, the sulphonate groups being located on opposite sides of the plane through the three middle atoms. The *trans* form of this ion was also found in the crystals of ammonium telluropentathionate⁹ and rubidium telluropentathionate hemitrihydrate.¹⁰

EXPERIMENTAL

Barium telluropentathionate trihydrate, $\text{BaTe}(\text{S}_2\text{O}_3)_2 \cdot 3\text{H}_2\text{O}$, was obtained by crystallization of crude barium telluropentathionate from dilute hydrochloric acid.¹¹ The crystals are light yellowish-green, flat prisms.

The intensity data used for the solution and first refinement of the structure were collected by the multiple-film technique. Integrated zero-layer and equi-inclination Weissenberg photographs were taken with $\text{CuK}\alpha$ radiation (Ni-filtered) for the $h0l$, $h1l$, $h2l$, and $0kl$ reflections. The intensities of the 1106 independent, observed reflections were measured visually.

The intensity data used for a more accurate refinement of the structure were measured on a Siemens automatic single-crystal diffractometer using $\text{MoK}\alpha$ radiation (Nb-filtered) and a scintillation counter.

The crystal was mounted with the b axis approximately parallel to the ϕ axis of the diffractometer, and the orientation of the crystal and preliminary unit cell dimensions were determined from the θ , χ , and ϕ angles of 10 non-coplanar reciprocal vectors.

The $\theta-2\theta$ scan technique and five-value procedure were used. The scan width was 0.70° for all reflections, and the maximum scan time per degree was 24 sec. For strong reflections the diffractometer automatically selects higher scan speed, and inserts a proper attenuation filter into the primary beam, if needed to avoid counting losses.

Two reflections, 700 and 0010, were used as reference, and measured two times each at intervals of 50 reflections. The intensities of the reference reflections decreased by 7 % during the data collection, and were used to bring the net intensities to a common scale.

Out of 2877 reflections attainable within $\theta = 28^\circ$, 488 were found to have net intensities below three times its standard deviation. These reflections were assigned an intensity equal to this limit, and labelled as unobserved.

Lorentz and polarization corrections were applied, and absorption corrections were carried out using a modified version of the Gaussian integration method described by Coppens *et al.*¹² The linear absorption coefficients for $\text{CuK}\alpha$ and $\text{MoK}\alpha$ radiations are 526 cm^{-1} and 66.2 cm^{-1} , respectively. The crystals used were prisms extended along the b axes and bounded by (100) and (001). The distances between the (100), between the (010), and between the (001) boundary faces were 0.036, 0.102, and 0.066 mm, respectively, for the crystal used to collect the $h0l-h2l$ film data, 0.014, 0.080, and 0.072 mm, respectively, for the crystal used to collect the $0kl$ film data, and 0.033, 0.167, and 0.110 mm, respectively, for the crystal used to collect the diffractometer data.

The film data were eventually also corrected for secondary extinction, using the formula $F_{\text{corr}} = KF_0(1 + \beta CI_0)$.¹³ Here $\beta = 2(1 + \cos^4 2\theta)/(1 + \cos^2 2\theta)^2$, and C was calculated to 1.34×10^{-7} .

The scattering factor curves used were those given in *International Tables for X-Ray Crystallography*,¹⁴ Table 3.3 1B for tellurium, Table 3.3. 1A for sulphur and oxygen, and the one given by Thomas and Umeda¹⁵ for barium ion. The curves for barium, tellurium and sulphur were corrected for anomalous dispersion using the $\Delta f'$ and $\Delta f''$ values given by Cromer,¹⁶ and taking the amplitude of f as the corrected value.

Least squares refinement was carried out with a program minimizing the function

$$r = \sum W(|F_o| - K|F_c|)^2$$

For the film data $W = [(3K)^2 + (2F_o)^2/4W_o + (0.07 F_o)^2]^{-1}$, where W_o is a weight factor based on the estimated reliability of the individual intensities. For the diffractometer data $W = 4(I_t - I_b)^2/F_o^2[I_t + I_b + k^2(I_t - I_b)^2]$ where I_t is the total intensity of a reflection, I_b is the background intensity, and k is the relative standard deviation in the scaling curve based on the reference reflections. Unobserved reflections were included with $|F_o|$ equal to the observable limit when $|F_c|$ exceeded this limit.

The programs used for calculations of setting angles and preparation of input tape data for the diffractometer, conversion of diffractometer output tape data to the computer, calculation of atomic distances and angles, and for secondary extinction corrections of the film data are written by K. Maartmann-Moe of this institute. The programs used for calculation of unit cell dimensions, film data processing, absorption corrections, two-dimensional Fourier summations, and least squares refinement were made available

by the Chemical Department of X-Ray Crystallography, Weizmann Institute of Science, Rehovoth, Israel. The programs used for diffractometer data processing and three-dimensional Fourier summations are written by K. Åse of this institute. The program used for drawing of illustrations is written by C. K. Johnson, Oak Ridge National Laboratory, Oak Ridge, Tennessee, USA. The calculations were carried out on an IBM 360/50H computer.

The unit cell dimensions, calculated by means of a least squares program using the θ angles ($\theta = 18 - 22^\circ$) of 10 reflections measured on the diffractometer are, $a = 11.139(3)$ Å, $b = 5.243(2)$ Å, $c = 21.306(6)$ Å, $\beta = 106.94(4)^\circ$. The space group is $P2_1/c$ (No. 14) with four BaTe(S₂O₃)₂·3H₂O formula units per unit cell.¹¹

THE STRUCTURE ANALYSIS

Two-dimensional Patterson maps, calculated on the basis of the $h0l$ and $0kl$ film data, revealed the positions of the two heavy atoms. The subsequent electron density maps allowed the location of the four sulphur atoms, and since two sulphur atoms are coordinated to the tellurium atom, it was possible to distinguish the tellurium atom from the barium ion. The sulphonate and water oxygen atoms were located in additional electron density maps.

Least squares refinement based on the 1106 independent, observed $h0l$, $h1l$, $h2l$ and $0kl$ reflections was then carried out. With individual isotropic thermal parameters for all atoms the R value ($(\sum ||F_o| - |F_c||) / \sum |F_o|$) converged at 0.120. Introduction of anisotropic thermal parameters for barium, tellurium and sulphur, followed by secondary extinction corrections, reduced the R value to 0.106 only.

None of the thermal parameters were unusually high, and the atomic distances and angles were not much different from those found in the rubidium salt of telluropentathionic acid.¹⁰ However, the standard deviations were rather high, the agreements between $|F_o|$ and $|F_c|$ for many of the strong

Table 1. Atomic coordinates for barium telluropentathionate trihydrate. Origin at a centre of symmetry. Standard deviations are given in parentheses.

	x	y	z
Ba	0.86098(2)	0.83134(6)	0.09725(1)
S(1)	0.16922(9)	0.62221(21)	0.04248(5)
S(2)	0.32916(10)	0.80347(23)	0.10351(5)
Te	0.40372(2)	0.50579(6)	0.19076(1)
S(4)	0.53083(10)	0.26496(23)	0.13903(5)
S(5)	0.70999(9)	0.41366(21)	0.18220(5)
O(1)	0.07601(31)	0.62342(91)	0.07634(18)
O(2)	0.13999(41)	0.77857(93)	-0.01489(18)
O(3)	0.20295(39)	0.36689(81)	0.03066(21)
O(4)	0.74671(29)	0.33338(69)	0.25033(14)
O(5)	0.78286(30)	0.29377(73)	0.14358(16)
O(6)	0.70578(31)	0.68819(64)	0.17507(16)
H ₂ O(1)	0.07150(32)	0.13818(69)	0.14962(17)
H ₂ O(2)	0.01078(33)	0.59785(79)	0.21127(17)
H ₂ O(3)	0.59990(38)	0.80036(106)	0.03228(19)

reflectrons were not good, and two-dimensional difference electron density maps showed unexpected high peaks.

The new set of data, recorded by means of a single-crystal diffractometer, was used for the further refinement of the structure. With atomic coordinates and thermal parameters as derived from the film data, the R value was 0.079. Refinement of the same parameters as in the film data set reduced the reliability index to 0.036, and by using anisotropic thermal parameters for all atoms, the refinement converged at $R=0.024$. Unobserved reflections are included when $|F_o|$ exceeds the observable limit. The final maximum shift of a parameter was about one fifth of its standard deviation.

A three-dimensional difference electron density map showed no peaks higher than $0.6 \text{ e}/\text{Å}^3$. The regions of positive electron density near the water oxygen atoms were too diffuse for location of the hydrogen atoms.

Tables 1 and 2 give the final atomic parameters with standard deviations from least squares. The structure factors, calculated on the basis of these parameters, and the observed ones, from the diffractometer data, are given in Table 3.

Table 2. Thermal parameters expressed in the form $\exp[-2\pi^2(h^2a^{-2}U_{11} + \dots + 2hka^{-1}b^{-1}U_{12} + \dots)]$. All values have been multiplied by 10^4 . Standard deviations are given in parentheses.

	U_{11}	U_{22}	U_{33}	U_{12}	U_{23}	U_{13}
Ba	247(1)	287(1)	302(1)	26(1)	64(1)	113(1)
S(1)	179(5)	201(5)	184(5)	-13(4)	0(4)	41(4)
S(2)	221(5)	240(6)	314(6)	-62(5)	18(5)	6(4)
Te	180(1)	308(2)	224(1)	3(1)	-28(1)	49(1)
S(4)	167(5)	295(6)	261(5)	-13(4)	-92(4)	43(4)
S(5)	152(5)	225(5)	206(4)	-10(4)	2(4)	58(4)
O(1)	262(19)	824(33)	467(22)	-159(21)	-225(22)	197(17)
O(2)	663(30)	755(33)	394(22)	-380(26)	376(23)	-207(20)
O(3)	523(27)	394(25)	740(30)	149(21)	-337(22)	-139(23)
O(4)	264(17)	400(20)	235(16)	-59(17)	70(15)	-7(13)
O(5)	252(17)	442(22)	447(20)	0(17)	-135(17)	180(16)
O(6)	355(20)	241(18)	443(20)	0(16)	39(16)	154(16)
H ₂ O(1)	362(20)	285(20)	458(21)	23(17)	58(17)	55(16)
H ₂ O(2)	345(21)	504(25)	397(20)	-82(18)	23(18)	58(16)
H ₂ O(3)	419(24)	1039(40)	417(23)	72(27)	60(25)	50(19)

RESULTS

The dimensions of the telluropentathionate ion, calculated from the atomic coordinates of Table 1, are listed in Table 4. The standard deviations given include estimated uncertainties in unit cell dimensions.

Fig. 1 is a view of the ion, as seen along a line through Te and the coordinate midpoint of S(2) and S(4), with principal bond lengths and angles. The largest differences between these values and those arrived at by the refinement based

Table 3. Observed and calculated structure factors ($\times 10$) for barium telluropentathionate trihydrate. A minus sign on $F(O)$ indicates an unobserved reflection.

F	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)
0	0	2	501	-460	4	0-22	447	-443	8	0-26	321	-303	12	0	4	416	-422	1	1	19	1009	-1003		
0	0	4	1755	-1748	4	0-20	418	-414	8	0-26	738	-722	12	0	6	591	-591	1	1	20	553	-520		
0	0	6	251	-211	4	0-18	437	-411	8	0-22	359	-375	12	0	8	-132	-104	1	1	21	805	-618		
0	0	8	1210	-1226	4	0-16	817	-314	8	0-20	-111	-96	13	0-18	499	-486	1	1	22	-121	-46			
0	0	10	2115	-2165	4	0-14	2623	2631	8	0-18	315	-318	13	0-16	547	-535	1	1	23	159	-159			
0	0	12	-157	138	4	0-12	1133	-1166	8	0-16	376	-317	13	0-14	-124	-123	1	1	24	425	421			
0	0	14	1736	-1746	4	0-10	1550	-1621	8	0-14	1766	-1696	13	0-12	1393	1415	1	1	25	111	-314			
0	0	16	740	-108	4	0-8	-150	67	8	0-12	-59	-27	13	0-10	523	-523	1	1	26	227	243			
0	0	18	1111	-1111	4	0-6	1439	-1415	8	0-10	1021	1123	13	0-8	652	-652	2	1-27	522	335				
0	0	20	1127	1154	4	0-4	2591	2602	8	0-8	1226	1191	13	0-6	347	359	2	1-26	478	400				
0	0	22	591	574	4	0-2	441	-450	8	0-6	242	228	13	0-4	567	-578	2	1-25	642	-627				
0	0	24	124	-157	4	0-0	2358	-2379	8	0-4	1737	-1739	13	0-2	1276	1288	2	1-24	652	-647				
0	0	26	-158	-21	4	0	2	1267	-1200	8	0-2	-111	-93	13	0	467	474	2	1-23	332	288			
1	0-24	366	-369	4	0	4	731	-712	8	0	0	560	498	13	0	2	1062	-1039	2	1-22	222	-208		
1	0-24	207	-205	4	0	6	2728	2710	8	0	2	-113	-37	13	0	4	-119	14	2	1-21	117	-113		
1	0-22	-174	-65	4	0	8	1455	1463	8	0	4	1038	1044	14	0-16	637	-681	2	1-20	672	-861			
1	0-22	130	127	4	0	10	2458	-2400	8	0	6	1268	-1407	14	0-14	235	-235	2	1-19	255	-193			
1	0-18	1065	1074	4	0	12	484	467	8	0	8	421	-337	14	0-12	-130	117	2	1-18	352	-352			
1	0-16	1392	-1365	4	0	14	-100	25	8	0	10	610	571	14	0-10	753	753	2	1-17	1499	1490			
1	0-14	1114	-1116	4	0	16	187	173	8	0	12	518	516	14	0-8	158	151	2	1-16	642	625			
1	0-12	413	415	4	0	18	854	801	8	0	14	923	922	14	0	16	815	-797	2	1-15	835	-739		
1	0-10	142	-165	4	0	20	1140	-1167	8	0	16	634	-660	14	0-6	315	315	2	1-14	753	752			
1	0-8	1627	1716	4	0	22	342	-326	8	0	18	368	-376	14	0-4	-211	71	2	1-13	441	-440			
1	0-6	610	-356	5	0-28	728	-704	9	0-26	312	-299	14	0	18	-159	52	2	1-12	674	-687				
1	0-4	811	-705	5	0-26	620	811	9	0-24	167	-148	14	0	1	646	-553	2	1-11	137	-125				
1	0-2	1131	1143	5	0-24	137	148	9	0-22	123	-132	14	0	2	1606	1555	2	1-10	1436	-1440				
1	0-2	124	-124	5	0-22	700	383	9	0-20	70	-70	14	0	4	1625	-1648	2	1-9	1415	-1415				
1	0-2	2566	2677	5	0-20	152	154	9	0-18	1216	1195	14	0	4	1996	1975	2	1-8	633	-610				
1	0	4	137	56	5	0-18	1340	-1346	9	0-16	829	-845	14	0	5	3113	-3193	2	1-7	2358	2329			
1	0	6	3363	-3443	5	0-16	818	824	9	0-14	470	448	14	0	6	-91	-9	2	1-6	639	673			
1	0	8	465	-457	5	0-14	875	886	9	0-12	565	-955	14	0	7	982	915	4	1-5	858	-863			
1	0	10	775	-741	5	0-12	1058	-844	9	0-10	1105	86	14	0	8	2222	-2189	4	1-4	2349	2369			
1	0	12	603	690	5	0-10	318	-330	9	0-8	2121	2194	14	0	9	1255	1297	2	1-3	1456	-1409			
1	0	14	1306	1223	5	0-8	2962	-2538	9	0-6	434	-607	14	0	10	452	-453	2	1-2	807	804			
1	0	16	1547	1529	5	0-6	819	715	9	0-4	622	-631	14	0	11	419	-453	2	1-1	700	620			
1	0	18	543	-547	5	0-4	2155	-2132	9	0-2	1275	-1234	14	0	12	1405	1463	2	1	0	1415	-1390		
1	0	20	107	258	5	0-2	153	145	9	0	0	351	-360	14	0	13	1366	1384	2	1	1	570	-567	
1	0	22	-111	-63	5	0	0	458	505	9	0	2	1908	1311	14	0	14	743	713	2	1	2	594	-579
1	0	24	1528	1063	5	0	2	1360	-1447	9	0	4	169	-196	14	0	15	1930	-1114	2	1	3	1529	2301
1	0	26	509	-601	5	0	4	667	616	9	0	6	751	-762	14	0	16	507	-921	2	1	4	1815	1760
1	0	28	465	-465	5	0	6	1250	1235	9	0	8	232	-249	14	0	17	323	-333	2	1	5	858	-862
2	0-26	293	-304	5	0	8	-103	101	9	0	10	234	-241	14	0	18	1138	-1146	2	1	6	2348	2073	
2	0-22	1461	-1461	5	0	10	604	558	9	0	12	834	802	14	0	19	1011	1915	2	1	7	947	-993	
2	0-20	458	466	5	0	12	1617	-1611	9	0	14	416	398	14	0	20	-139	-53	2	1	8	334	-320	
2	0-18	-111	89	5	0	14	854	-838	9	0	16	230	-247	14	0	21	491	-463	2	1	9	144	-140	
2	0-16	250	-261	5	0	16	450	-458	9	0	18	62	-62	14	0	22	235	245	2	1	10	1415	-1474	
2	0-14	844	828	5	0	18	206	-291	10	0-22	-113	-37	14	0	23	179	185	2	1	11	517	-601		
2	0-12	2113	-2173	5	0	20	135	79	10	0-20	461	-448	14	0	24	243	222	2	1	12	557	-670		
2	0-10	557	551	5	0	22	338	-326	10	0-18	-116	125	14	0	25	266	-264	2	1	13	632	640		
2	0-8	2474	2450	6	0-26	587	-575	10	0-16	1557	1331	14	0	26	-123	43	2	1	14	493	485			
2	0-6	111	110	6	0-14	615	-620	10	0-14	623	-618	14	0	27	159	-159	2	1	15	1415	-1415			
2	0-4	348	344	6	0-12	1110	554	10	0-12	472	-478	14	0	28	276	303	2	1	16	710	710			
2	0-2	2612	-2552	6	0-10	466	-472	10	0-10	175	183	14	0	29	-114	-112	2	1	17	1398	-1333			
2	0	519	-632	6	0-8	627	615	10	0-8	1069	-1025	14	0	30	398	-297	2	1	18	323	304			
2	0	2	1613	1650	6	0-6	1625	-1621	10	0-6	1550	1562	14	0	31	675	-696	2	1	19	121	143		
2	0	4	178	-185	6	0-4	365	375	10	0-4	367	-368	14	0	32	372	-372	2	1	20	675	-746		
2	0	6	-53	-14	6	0-2	1530	1563	10	0-2	1129	-1137	14	0	33	711	713	2	1	21	179	-179		
2	0	8	2716	-2742	6	0-0	688	-685	10	0	0	-190	275	14	0	34	153	153	2	1	22	437	-454	
2	0	10	658	-763	6	0-8	313	296	10	0	2	950	-849	14	0	35	253	267	2	1	23	175	185	
2	0	12	1823	1818	6	0-6	1166	-1237	10	0	4	1275	1457	14	0	36	444	424	2	1	24	251	269	
2	0	14	1065	1060	6	0-4	1118	-1120	10	0	6	1757	1901	14	0	37	-111	11	2	1	25	498	489	
2	0	16	346	317	6	0-2	3450	3740	10	0	8	1128	-1142	14	0	38	656	689	3	1-27	151	-163		
2	0	18	777	-778	6	0	0	-140	-42	10	0	10	250	235	14	0	39	435	474	3	1-26	211	205	
2	0	20	467	-446	6	0	2	714	-711	10	0	12	332	-309	14	0	40	-93	27	3	1-25	896	820	
2	0	22	1755	1764	6	0	4	1668	-1672	10	0	14	171	161	14	0	41	1767	-1103	3	1-24	399	-393	
2	0	24	-113	6	6	0	6	1251	-1252	11	0-24	571	569	14	0	42	372	-372	3	1-23	444	-442		
2	0	26	751	724	6	0	8	1568	1585	11	0-22	-151	-120	14	0	43	557	561	3	1-22	-134	-40		
2	0	28	257	261	6	0	10	412	404	11	0-20	661	-677	14	0	44	794	-784	3	1-21	394	-310		
2	0	30	617	-955	6	0	12	1102	-1166	11	0-18	376	367	14	0	45	815	931	3	1-20	718	-736		
2	0	32	186	-175	6	0	14	132	-158	11	0-16	252	310											

Table 3. Continued.

F	K	L	F(0)	F(C)	H	K	L	F(0)	F(C)	H	K	L	F(0)	F(C)	H	K	L	F(0)	F(C)	H	K	L	F(0)	F(C)
2	1	8	457	475	5	1	9	148	-155	7	1	15	-115	-28	13	1	-12	517	497	13	1	-6	693	-684
3	1	5	1566	-1586	5	1	10	660	555	7	1	16	312	-341	13	1	-11	1320	-1339	13	1	-5	-123	81
3	1	10	278	388	5	1	11	615	-555	7	1	17	725	-751	13	1	-10	322	-301	13	1	-4	696	-684
3	1	11	355	-355	5	1	12	145	146	7	1	18	554	-543	13	1	-9	453	451	13	1	-3	530	513
3	1	12	465	-477	5	1	13	514	511	7	1	19	532	545	10	1	-8	579	-593	13	1	-2	181	194
3	1	13	684	689	5	1	14	355	-352	8	1	-26	481	496	13	1	-7	591	617	13	1	-1	536	-530
3	1	14	165	154	5	1	15	265	-249	8	1	-25	214	214	13	1	-6	-113	34	13	1	0	718	694
3	1	15	165	-157	5	1	16	128	-106	8	1	-24	-111	-6	13	1	-5	354	-354	13	1	1	529	524
3	1	16	732	795	5	1	17	523	521	8	1	-23	277	-284	13	1	-4	752	749	13	1	2	242	221
3	1	17	735	722	5	1	18	121	-72	8	1	-22	517	-204	13	1	-3	1136	1152	13	1	3	442	-454
3	1	18	510	467	5	1	19	392	373	8	1	-21	-164	35	13	1	-2	333	337	13	1	4	176	-172
3	1	19	434	-475	5	1	20	153	106	8	1	-20	273	-264	13	1	-1	1009	-1226	14	1	-15	332	-340
3	1	20	-110	-112	5	1	21	523	-545	8	1	-19	950	941	13	1	0	594	-574	14	1	-14	431	-444
3	1	21	250	-255	5	1	22	163	123	8	1	-18	566	595	13	1	1	181	-185	14	1	-13	-126	-84
3	1	22	167	-161	6	1	-27	-118	-106	8	1	-17	336	-356	10	1	2	690	-569	14	1	-12	161	-151
3	1	23	564	551	6	1	-26	253	-274	8	1	-16	614	619	13	1	3	892	739	14	1	-11	480	494
3	1	24	147	-131	6	1	-25	158	162	8	1	-15	-161	13	10	1	4	161	-155	14	1	-10	-126	106
4	1	-27	485	454	6	1	-24	370	-319	8	1	-14	456	-438	13	1	5	718	-719	14	1	-9	433	-431
4	1	-24	857	-815	6	1	-23	579	-580	8	1	-13	224	-217	13	1	6	334	354	14	1	-8	368	354
4	1	-21	-112	67	6	1	-22	133	-116	8	1	-12	744	-754	13	1	7	424	464	14	1	-7	-122	108
4	1	-24	277	275	6	1	-21	944	953	8	1	-11	534	-527	13	1	8	533	663	14	1	-6	242	230
4	1	-22	170	157	6	1	-20	594	603	8	1	-10	353	-390	10	1	9	336	-324	14	1	-5	235	-225
4	1	-22	424	411	6	1	-19	452	-453	8	1	-9	378	370	13	1	10	-131	146	14	1	-4	239	-167
4	1	-21	164	-165	6	1	-18	945	947	8	1	-8	524	517	13	1	11	637	641	14	1	-3	423	-423
4	1	-20	427	-434	6	1	-17	302	-301	8	1	-7	542	-610	13	1	12	351	-325	14	1	-2	293	-186
4	1	-19	1185	-1232	6	1	-16	223	-225	8	1	-6	942	944	13	1	13	720	733	14	1	-1	674	671
4	1	-18	614	-635	6	1	-15	1047	1054	8	1	-5	357	-310	11	1	-2	333	-377	14	1	0	667	613
4	1	-17	511	487	6	1	-14	841	-845	8	1	-4	558	891	11	1	-2	630	627	14	1	1	895	925
4	1	-14	572	-577	6	1	-13	573	-571	8	1	-3	-115	81	11	1	-1	114	-115	14	1	2	114	-114
4	1	-11	-35	70	6	1	-12	-56	-5	8	1	-2	753	-724	11	1	-2	193	-185	14	1	3	1203	1244
4	1	-14	-52	-31	6	1	-11	1251	1304	8	1	-1	546	-615	11	1	-1	446	-451	14	1	4	1216	1246
4	1	-11	124	-146	6	1	-10	955	985	8	1	0	844	-849	11	1	-10	663	-663	14	1	5	-93	-123
4	1	-12	1221	1249	6	1	-9	551	573	8	1	1	1451	1476	11	1	-17	-115	-334	14	1	6	1383	1464
4	1	-10	742	-742	6	1	-8	1658	6808	8	1	2	274	-278	11	1	-16	279	-317	14	1	7	523	-543
4	1	-10	1140	1112	6	1	-7	2134	-2115	8	1	3	156	175	11	1	-15	796	777	14	1	8	2343	-1492
4	1	-5	1642	-1702	6	1	-6	722	-736	8	1	4	1154	1167	11	1	-14	755	749	14	1	9	608	-645
4	1	-1	664	-675	6	1	-5	654	643	8	1	5	553	-619	11	1	-13	387	-378	14	1	10	635	626
4	1	-7	494	675	6	1	-4	1226	-1238	8	1	6	252	264	11	1	-12	795	724	14	1	11	637	620
4	1	-4	154	-154	6	1	-3	344	-347	8	1	7	274	-274	11	1	-11	45	-45	14	1	12	523	-543
4	1	-5	1199	1420	6	1	-2	672	-648	8	1	8	647	-658	11	1	-10	-115	-85	14	1	13	1314	1340
4	1	-4	527	676	6	1	-1	154	-85	8	1	9	646	-641	11	1	-9	271	-264	14	1	14	226	-224
4	1	-3	404	-432	6	1	0	549	547	8	1	10	465	-468	11	1	-8	643	-642	14	1	15	-130	-10
4	1	-7	226	2345	6	1	1	1678	1665	8	1	11	753	713	11	1	-7	-117	-107	14	1	16	1396	1364
4	1	-11	1163	1171	6	1	2	584	584	8	1	12	-119	73	11	1	-6	-128	-119	14	1	17	519	-519
4	1	0	120	395	6	1	3	1966	-1858	8	1	13	-113	27	11	1	-5	1079	1073	14	1	18	474	-474
4	1	1	1854	-1854	6	1	4	201	-181	8	1	14	252	231	11	1	-4	-114	15	14	1	19	308	-376
4	1	2	1476	-1458	6	1	5	234	217	8	1	15	565	-511	11	1	-3	376	-291	14	1	20	233	-216
4	1	5	125	-123	6	1	6	1214	-1195	8	1	16	368	325	11	1	-2	574	568	14	1	21	1306	1266
4	1	4	1268	-1113	6	1	7	585	585	8	1	17	61	183	11	1	-1	128	-139	14	1	22	-108	-41
4	1	1	1574	1560	6	1	8	645	-550	9	1	-26	380	363	11	1	0	-113	62	14	1	23	563	555
4	1	6	165	-165	6	1	9	142	-165	9	1	-25	320	-355	11	1	1	-119	-113	14	1	24	232	-220
4	1	7	1335	-1293	6	1	10	659	652	9	1	-24	227	222	11	1	2	917	-937	14	1	25	-129	50
4	1	10	125	-123	6	1	11	1258	1241	9	1	-23	268	-287	11	1	3	823	-823	14	1	26	268	-268
4	1	9	130	124	6	1	12	755	757	9	1	-22	430	-441	11	1	4	551	-545	14	1	27	180	192
4	1	10	580	591	6	1	13	677	-665	9	1	-21	450	487	11	1	5	772	773	14	1	28	474	493
4	1	11	528	-520	6	1	14	121	-115	9	1	-20	658	-675	11	1	6	588	603	14	1	29	523	-515
4	1	12	276	-252	6	1	15	391	-151	9	1	-19	455	-507	11	1	7	-123	-122	14	1	30	190	213
4	1	13	1332	-1324	6	1	16	1254	-1248	9	1	-18	423	-441	11	1	8	542	-535	14	1	31	158	-156
4	1	14	890	-875	6	1	17	571	572	9	1	-17	1047	1447	11	1	9	310	-337	14	1	32	847	-834
4	1	15	1291	1367	6	1	18	303	-339	9	1	-16	766	734	11	1	10	-129	91	14	1	33	149	135
4	1	16	-105	8	6	1	19	222	-311	9	1	-15	270	-280	11	1	11	298	311	14	1	34	363	366
4	1	17	-105	-54	6	1	20	244	275	9	1	-14	576	568	11	1	-12	275	-285	14	1	35	493	514
4	1	18	154	147	6	1	21	154	-164	9	1	-13	425	-441	11	1	-11	1476	-1476	14	1	36	191	-184
4	1	19	-114	23	7	1	-26	-116	-71	9	1	-12	761	-772	12	1	-10	473	-466	14	1	37	496	500
4	1	20	523	563	7	1	-25	683	-667	9	1	-11	729	725	12	1	-9	-115	-41	14	1	38	633	604
4	1	21	223	244	7	1	-24	597	-531	9	1	-10	1313	-1325	12	1	-8	528	515	14	1	39	376	-387
4	1	22	253	-377	7	1	-23	578	563	9	1	-9	565	-505	12</									

Table 3. Continued.

H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)
2	2-22	211	204	6	2-22	632	-622	6	2-15	932	-740	8	2-2	625	595	11	2-12	405	405
2	2-22	158	135	4	2-21	634	653	6	2-14	143	-141	8	2-1	1123	-1117	11	2-11	172	159
2	2-24	201	352	4	2-20	666	651	6	2-13	666	-641	8	2-2	223	-194	11	2-10	-112	19
2	2-22	-112	103	4	2-19	-105	19	6	2-12	326	300	8	2-1	374	-381	11	2-9	351	-366
2	2-22	231	-252	4	2-18	664	657	6	2-11	685	686	8	2-2	908	-900	11	2-8	695	714
2	2-21	481	-474	4	2-17	567	-980	6	2-10	1574	-1564	8	2-3	604	713	11	2-7	637	-658
2	2-20	163	777	4	2-16	1343	-1355	6	2-9	1221	1215	8	2-4	481	473	11	2-6	767	-773
2	2-15	725	-740	4	2-15	-112	-75	6	2-8	1014	1356	8	2-5	372	334	11	2-5	-110	-106
2	2-18	779	-779	4	2-14	176	181	6	2-7	375	315	8	2-6	331	-335	11	2-4	590	-568
2	2-17	645	636	4	2-13	376	376	6	2-6	601	577	8	2-7	-114	-39	11	2-3	1054	1000
2	2-16	363	-403	4	2-12	376	-378	6	2-5	320	-315	8	2-8	459	464	11	2-2	964	970
2	2-15	1019	1015	4	2-11	612	625	6	2-4	441	-420	8	2-9	215	-225	11	2-1	489	488
2	2-14	635	513	4	2-10	572	558	6	2-3	1011	-1038	8	2-10	330	332	11	2-0	195	-187
2	2-13	458	-454	4	2-9	553	-564	6	2-2	222	697	8	2-15	453	453	11	2-0	195	-187
2	2-12	653	-579	4	2-8	1015	1119	6	2-1	393	384	8	2-16	130	143	11	2-0	195	-187
2	2-11	526	-511	4	2-7	1369	-1375	6	2-0	1508	-1521	8	2-12	1148	-1156	11	2-2	311	287
2	2-10	1121	1044	4	2-6	1444	-1438	6	2-1	561	958	8	2-13	624	533	11	2-2	311	287
2	2-9	1069	-1064	4	2-5	612	-608	6	2-2	722	697	8	2-14	271	269	11	2-4	-117	-10
2	2-8	1234	-1214	4	2-4	375	375	6	2-3	327	333	8	2-15	130	143	11	2-6	596	-593
2	2-7	1537	-1534	4	2-3	1168	1178	6	2-4	1258	1255	8	2-16	886	893	11	2-7	370	392
2	2-6	1754	1742	4	2-2	231	-214	6	2-5	1013	-1020	9	2-23	-115	1	11	2-8	411	443
2	2-5	2037	2029	4	2-1	671	659	6	2-6	582	-367	9	2-22	259	-267	11	2-9	152	122
2	2-4	2377	-2369	4	2-0	328	-316	6	2-7	950	-386	9	2-21	476	-395	12	2-10	426	-436
2	2-3	277	-269	4	2-1	151	146	6	2-8	464	431	9	2-20	146	-127	12	2-10	755	747
2	2-2	475	-467	4	2-2	2533	2544	6	2-9	175	-166	9	2-19	256	-262	12	2-19	167	-159
2	2-1	517	-513	4	2-3	384	-381	6	2-10	752	-720	9	2-18	130	-122	12	2-18	237	-230
2	2-0	510	510	4	2-4	502	-494	6	2-11	742	749	9	2-17	650	653	12	2-17	425	-410
2	2-1	2311	-2283	4	2-5	594	-584	6	2-12	244	-212	9	2-16	527	-543	12	2-16	911	-914
2	2-2	337	-339	4	2-6	442	-438	6	2-13	413	406	9	2-15	723	715	12	2-15	421	-425
2	2-3	372	367	4	2-7	608	605	6	2-14	936	913	9	2-14	1048	1045	12	2-14	244	242
2	2-4	1730	-1707	4	2-8	437	435	6	2-15	209	-217	9	2-13	136	133	12	2-13	174	170
2	2-5	1791	1733	4	2-9	1022	1123	6	2-16	274	-274	9	2-12	154	-155	12	2-12	710	686
2	2-6	1254	1153	4	2-10	-112	-113	6	2-17	162	-174	9	2-11	545	-541	12	2-11	538	511
2	2-7	659	649	4	2-11	478	-476	6	2-18	-121	92	9	2-10	193	-141	12	2-10	454	450
2	2-8	331	324	4	2-12	1156	1169	6	2-19	-116	17	9	2-9	945	-939	12	2-9	-123	80
2	2-9	356	-348	4	2-13	851	-638	6	2-20	234	-227	9	2-8	201	214	12	2-8	-117	-16
2	2-10	215	156	4	2-14	426	-465	7	2-25	531	516	9	2-7	276	265	12	2-7	112	46
2	2-11	582	-577	4	2-15	742	-714	7	2-22	256	-251	9	2-6	276	265	12	2-6	-112	23
2	2-12	245	206	4	2-16	1165	-696	7	2-24	-116	51	9	2-5	741	739	12	2-5	380	-391
2	2-13	317	-282	4	2-17	443	446	7	2-23	724	-749	9	2-4	1731	1025	12	2-4	263	245
2	2-14	1567	-1558	4	2-18	564	551	7	2-22	560	-577	9	2-3	313	-302	12	2-3	156	167
2	2-15	675	372	4	2-19	459	459	7	2-21	113	-55	9	2-2	733	723	12	2-2	911	911
2	2-16	494	451	4	2-20	442	-449	7	2-20	413	406	9	2-1	636	-577	12	2-1	668	687
2	2-17	311	195	4	2-21	-117	32	7	2-19	451	442	9	2-0	796	-763	12	2-0	417	410
2	2-18	505	-511	4	2-22	461	443	7	2-18	411	-311	9	2-1	430	-433	12	2-1	-117	27
2	2-19	458	-531	4	2-23	151	158	7	2-17	321	314	9	2-2	239	201	12	2-2	455	468
2	2-20	232	-234	5	2-25	124	-129	7	2-16	454	951	9	2-3	452	455	12	2-3	-112	46
2	2-21	572	-562	5	2-24	724	-724	7	2-15	374	-374	9	2-4	542	-544	12	2-4	414	-425
2	2-22	-111	1177	5	2-23	591	510	7	2-14	575	503	9	2-5	878	807	12	2-5	374	-388
2	2-23	205	-154	5	2-22	205	222	7	2-13	865	-385	9	2-6	435	427	12	2-6	331	353
2	2-24	458	-503	5	2-21	489	477	7	2-12	1615	-1631	9	2-7	145	170	13	2-17	-122	-50
2	2-25	310	-325	5	2-20	453	446	7	2-11	343	344	9	2-8	323	331	13	2-18	414	-404
2	2-26	273	266	5	2-19	164	174	7	2-10	172	329	9	2-9	329	-364	13	2-15	416	-404
2	2-27	247	216	5	2-18	463	-367	7	2-9	1167	1172	9	2-10	655	-689	13	2-14	776	-799
2	2-28	316	301	5	2-17	647	-254	7	2-8	-102	-56	9	2-11	268	-222	13	2-13	-126	-137
2	2-29	1733	1732	5	2-16	610	758	7	2-7	763	772	9	2-12	328	335	13	2-12	166	205
2	2-30	-135	15	5	2-15	1258	-1267	7	2-6	-101	18	9	2-13	-123	-11	13	2-11	565	557
2	2-31	65	65	5	2-14	672	-672	7	2-5	642	-623	9	2-14	219	-223	13	2-10	425	-423
2	2-32	476	-477	5	2-13	570	46	7	2-4	637	844	13	2-15	352	344	13	2-9	538	513
2	2-33	291	-267	5	2-12	642	-628	7	2-3	1042	-1049	13	2-16	422	599	13	2-8	202	170
2	2-34	-134	52	5	2-11	911	907	7	2-2	1358	-1353	13	2-17	-115	102	13	2-7	-116	10
2	2-35	227	224	5	2-10	1437	1390	7	2-1	514	-517	13	2-18	315	315	13	2-6	141	150
2	2-36	1025	-1023	5	2-9	1116	1116	7	2-0	1338	140	13	2-19	431	-433	13	2-5	446	-431
2	2-37	651	-644	5	2-8	363	-364	7	2-1	638	635	13	2-20	657	-649	13	2-4	630	-631
2	2-38	567	517	5	2-7	746	-745	7	2-2	-110	-64	13	2-21	124	-91	13	2-3	439	-430
2	2-39	1436	1435	5	2-6	421	435	7	2-3	475	510	13	2-22	351	349	13	2-2	237	230
2	2-40	-62	-61	5	2-5	351	-357	7	2-4	-432	-395	13	2-23	351	430	13	2-1	292	230
2	2-41	1111	-1105	5	2-4	421	-421	7	2-5	428	-434	13	2-24	449	-444	13	2-0	233	-203
2	2-42	555	-664	5	2-3	423	494	7	2-6	1344	1352	13	2-25	249	305	13	2-1	235	289
2	2-43	547	-561	5	2-2	851	-829	7	2-7	658	-658	13	2-26	410	425	13	2-2	162	-174
2	2-44	726	-625	5	2-1	1279	1311	7	2-8	860	-841	13	2-27	230	-213	14	2-11	-126	-15
2	2-45	1041	1015	5	2-0	1581	1561	7	2-9	351	-343	13	2-28	677	684	14	2-10	248	-240
2	2-46	1414	-1423	5	2-1	412	155	7	2-10	511	-518	13	2-29	856	-859	14	2-9	421	441
2	2-47	1471	-1448	5	2-2	751	-752	7	2-11	455	453	13	2-30	885	-872	14	2-8	342	363
2	2-48	637	651	5	2-3	-105	-18	7	2-12	523	623	13	2-31	509	-594	14	2-7	121	65
2	2-49	1257	1323	5	2-4	277	295	7	2-13	621									

Table 3. Continued.

F	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)				
1	3-21	-108	51	3	3-14	-104	-116	5	3-4	535	533	7	3-12	-115	-72	10	3	5	827	-864								
1	3-20	319	-222	3	3-13	537	935	5	3-3	-169	33	7	3-13	323	325	10	3	6	-119	-78								
1	3-15	485	-474	3	3-12	1137	100	5	3-2	431	-44	7	3-14	-127	-135	10	3	7	444	464								
1	3-16	-114	-2	3	3-11	485	490	5	3-1	670	-674	7	3-15	615	511	10	3	8	-122	-38								
1	3-17	635	-441	3	3-10	227	215	5	3	150	-152	7	3-16	-122	-67	10	3	9	436	403								
1	3-14	466	464	3	3-9	1156	1150	5	3	1031	986	8	3-23	-115	0	10	3	10	-130	-110								
1	3-15	1227	1247	3	3-8	307	395	5	3	2	194	203	8	3-22	284	293	11	3	-19	249	226							
1	3-14	319	-378	3	3-7	1744	-1756	5	3	3	1959	1378	8	3-21	374	-373	11	3	-18	230	-512							
1	3-13	430	-415	3	3-6	347	-337	5	3	4	232	-251	8	3-20	-110	-13	11	3	-17	377	-374							
1	3-12	214	-227	3	3-5	460	-470	5	3	5	1142	-1134	8	3-19	219	-213	11	3	-16	-119	113							
1	3-11	153	156	3	3-4	179	-167	5	3	6	274	259	8	3-18	314	-333	11	3	-15	171	-174							
1	3-10	-96	-12	3	3-3	673	873	5	3	7	124	-128	8	3-17	891	-795	11	3	-14	273	-262							
1	3-9	556	-555	3	3-2	629	391	5	3	8	-110	-78	8	3-16	128	-95	11	3	-13	687	-670							
1	3-8	511	-513	3	3-1	171	-173	5	3	9	138	147	8	3-15	1010	1323	11	3	-12	311	311							
1	3-7	1208	-1254	3	3	0	-116	-131	5	3	10	135	129	8	3-14	231	-224	11	3	-11	931	852						
1	3-6	252	-264	3	3	1	761	766	5	3	11	255	-232	8	3-13	443	483	11	3	-10	173	152						
1	3-5	2367	2363	3	3	2	451	-425	5	3	12	446	444	8	3-12	256	-273	11	3	-9	646	626						
1	3-4	-95	-26	3	3	3	1346	-1745	5	3	13	1017	1016	8	3-11	623	-613	11	3	-8	239	-192						
1	3-3	560	555	3	3	4	203	182	5	3	14	215	-213	8	3-10	-106	-35	11	3	-7	843	-842						
1	3-2	113	-55	3	3	5	1372	-1374	5	3	15	614	-643	8	3-9	298	293	11	3	-6	147	140						
1	3-1	444	-431	3	3	6	173	-147	5	3	16	135	-153	8	3-8	-112	-97	11	3	-5	-123	-62						
1	3	0	266	253	3	3	7	1363	1030	5	3	17	463	-472	8	3-7	869	-361	11	3	-4	132	-187					
1	3	1	254	254	3	3	8	104	104	5	3	18	243	-243	8	3-6	646	646	11	3	-3	169	-169					
1	3	2	261	-258	3	3	9	-130	69	5	3	19	-117	65	8	3-5	559	653	11	3	-2	332	-295					
1	3	3	1726	-1742	3	3	10	-106	-15	5	3	20	424	-422	8	3-4	-114	-27	11	3	-1	179	561					
1	3	4	42	418	3	3	11	1765	1352	5	3	21	-119	-25	8	3-3	-119	1292	11	3	0	-520	9					
1	3	5	655	661	3	3	12	126	-70	5	3	22	771	-754	8	3-2	270	-257	11	3	1	1445	1033					
1	3	6	148	145	3	3	13	-103	-43	5	3	23	-114	-44	8	3-1	914	-914	11	3	2	236	-236					
1	3	7	404	421	3	3	14	167	165	5	3	24	-102	11	8	3	0	825	921	11	3	3	594	-500				
1	3	8	645	-458	3	3	15	952	-943	5	3	25	166	158	8	3	1	303	-311	11	3	4	222	228				
1	3	9	736	-750	3	3	16	-112	72	5	3	26	-166	46	8	3	2	-117	5	11	3	5	-124	-104				
1	3	10	-49	11	3	3	17	535	545	5	3	27	155	110	8	3	3	954	-379	11	3	6	-116	-34				
1	3	11	-10	73	3	3	18	-115	144	5	3	28	-115	-63	8	3	4	433	-453	11	3	7	243	-212				
1	3	12	271	282	3	3	19	-113	71	5	3	29	-115	-63	8	3	5	-116	31	12	3	-17	495	322				
1	3	13	115	-1153	3	3	20	270	190	5	3	30	753	744	8	3	6	166	173	12	3	-16	-133	160				
1	3	14	331	-277	3	3	21	-121	23	5	3	31	428	424	8	3	7	1217	1233	12	3	-15	789	-789				
1	3	15	115	102	3	3	22	236	-191	5	3	32	1413	-1401	8	3	8	152	-153	12	3	-14	232	213				
1	3	16	-107	165	3	3	23	156	-724	5	3	33	-116	-84	8	3	9	537	-543	12	3	-13	591	-591				
1	3	17	1070	1074	3	3	24	156	-182	5	3	34	688	-676	8	3	10	-142	-55	12	3	-12	331	-330				
1	3	18	154	-141	3	3	25	245	-246	5	3	35	479	-463	8	3	11	225	-243	12	3	-11	-123	-27				
1	3	19	239	-237	3	3	26	-103	21	5	3	36	-99	172	8	3	12	-120	33	12	3	-10	-119	70				
1	3	20	122	76	3	3	27	657	660	5	3	37	-104	-48	8	3	13	-123	-58	12	3	-9	-129	113				
1	3	21	153	165	3	3	28	421	-222	5	3	38	362	358	8	3	14	866	-136	12	3	-8	-124	-34				
1	3	22	225	217	3	3	29	119	27	5	3	39	-349	-346	8	3	15	22	146	-174	12	3	-7	716	723			
1	3	23	431	-250	3	3	30	-108	64	5	3	40	-1076	1686	9	3	16	451	445	12	3	-6	337	-323				
1	3	24	269	-322	3	3	31	603	610	5	3	41	-117	-175	9	3	17	-119	91	12	3	-5	695	-673				
1	3	25	-114	15	3	3	32	-108	-44	5	3	42	113	-1040	537	9	3	18	537	-337	12	3	-4	169	-27			
1	3	26	845	848	3	3	33	560	-614	5	3	43	-132	130	9	3	19	-114	-53	12	3	-3	175	-173				
1	3	27	-127	47	3	3	34	139	99	5	3	44	829	-818	9	3	20	24	-254	12	3	-2	-120	-43				
1	3	28	-137	33	3	3	35	1166	-1152	5	3	45	-116	-57	9	3	21	352	-373	12	3	-1	293	277				
1	3	29	316	343	3	3	36	147	147	5	3	46	736	732	9	3	22	324	303	12	3	0	-126	115				
1	3	30	-115	160	3	3	37	-104	-44	5	3	47	-117	108	9	3	23	-116	-110	12	3	1	460	-455				
1	3	31	210	-154	3	3	38	239	-246	5	3	48	-123	65	9	3	24	592	589	12	3	2	-122	4				
1	3	32	143	-184	3	3	39	243	256	5	3	49	-116	-41	9	3	25	-114	83	12	3	3	814	814				
1	3	33	412	-389	3	3	40	126	171	5	3	50	1302	1293	9	3	26	-111	912	133	12	3	-12	-125	-56			
1	3	34	747	-749	3	3	41	-99	74	5	3	51	123	106	9	3	27	276	214	13	3	-11	521	-521				
1	3	35	145	145	3	3	42	771	771	5	3	52	-116	-48	9	3	28	1267	-1267	13	3	-10	169	-169				
1	3	36	1485	1487	3	3	43	683	-580	5	3	53	165	155	9	3	29	-116	-62	13	3	-9	697	708				
1	3	37	366	-252	3	3	44	141	-153	5	3	54	1610	-1393	9	3	30	-7	335	-333	13	3	-8	-125	-82			
1	3	38	1117	1124	3	3	45	2105	-2144	5	3	55	-117	75	9	3	31	157	-189	13	3	-7	329	302				
1	3	39	-95	32	3	3	46	-116	-116	5	3	56	111	262	257	9	3	32	319	-319	13	3	-6	-119	9			
1	3	40	115	115	3	3	47	1134	-1134	5	3	57	1413	-1413	9	3	33	515	509	13	3	-5	236	-236				
1	3	41	221	216	3	3	48	161	-161	5	3	58	-114	-94	9	3	34	-119	17	13	3	-4	165	163				
1	3	42	211	-122	3	3	49	1169	-1169	5	3	59	-119	113	9	3	35	-121	56	13	3	-3	174	-193				
1	3	43	535	-523	3	3	50	209	-205	5	3	60																

Table 3. Continued.

H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)
1	4	-8	613	-623	3	4	12	646	-684	6	4	-8	566	990	9	4	-8	-117	-16	1	5	3	619	-595
1	4	-4	269	-265	3	4	13	336	306	6	4	-7	221	-222	9	4	-7	222	-324	1	5	4	786	-768
1	4	-3	527	525	3	4	14	152	-157	6	4	-6	108	-25	9	4	-6	506	-573	1	5	5	139	122
1	4	-2	519	908	3	4	15	207	-227	6	4	-5	742	757	9	4	-5	611	-369	1	5	6	166	-182
1	4	-1	462	466	3	4	16	169	-178	6	4	-4	625	-637	9	4	-4	695	-691	1	5	7	172	-154
1	4	0	1439	-1120	3	4	17	217	-192	6	4	-3	559	614	9	4	-3	261	-218	1	5	8	436	404
1	4	1	-111	72	3	4	18	-113	63	6	4	-2	161	173	9	4	-2	239	202	1	5	9	633	-685
1	4	2	161	-113	3	4	19	216	-203	6	4	-1	266	-118	9	4	-1	767	744	1	5	10	601	640
1	4	3	274	-222	4	4	-23	479	-482	6	4	0	783	-792	9	4	0	-125	-92	1	5	11	531	514
1	4	4	574	557	4	4	-22	211	-232	5	4	1	579	-536	9	4	1	371	353	1	5	12	224	-229
1	4	5	871	-652	4	4	-21	176	-171	6	4	2	301	310	9	4	2	469	457	1	5	13	-117	-91
1	4	6	226	-753	4	4	-20	222	222	6	4	3	567	-574	9	4	3	472	-416	1	5	14	390	-345
1	4	7	231	249	4	4	-19	275	283	6	4	4	253	239	9	4	4	690	-711	1	5	15	-118	85
1	4	8	1134	1123	4	4	-18	200	215	6	4	5	655	549	9	4	5	299	-277	1	5	16	-115	35
1	4	9	755	783	4	4	-17	635	483	6	4	6	400	-401	9	4	6	164	187	1	5	17	123	162
1	4	10	-116	3	4	4	-16	871	-573	6	4	7	627	672	9	4	7	-134	-125	1	5	18	-125	129
1	4	11	-125	-11	4	4	-15	258	335	6	4	8	657	456	9	4	8	363	381	2	5	19	-115	45
1	4	12	495	-56	4	4	-14	273	304	6	4	9	185	-173	9	4	9	231	267	2	5	18	124	134
1	4	13	-115	51	4	4	-13	659	-708	6	4	10	416	-413	10	4	-17	131	127	2	5	17	508	528
1	4	14	331	607	4	4	-12	353	-323	6	4	11	401	-394	10	4	-16	164	181	2	5	16	275	-273
1	4	15	390	-368	4	4	-11	74	-754	6	4	12	350	407	10	4	-15	253	-247	2	5	15	493	-531
1	4	16	151	-150	4	4	-10	-112	58	6	4	13	233	-232	10	4	-14	493	-601	2	5	14	258	-244
1	4	17	139	-130	4	4	-9	523	531	6	4	14	662	637	10	4	-13	479	-481	2	5	13	-107	37
1	4	18	313	300	4	4	-8	1241	1236	6	4	15	-149	155	10	4	-12	-116	24	2	5	12	-197	124
1	4	19	365	375	4	4	-7	493	500	7	4	-21	-123	131	10	4	-11	157	179	2	5	11	-115	-14
1	4	20	222	-233	4	4	-6	758	-759	7	4	-20	325	344	10	4	-10	561	552	2	5	10	675	673
1	4	21	113	114	4	4	-5	397	374	7	4	-19	619	-623	10	4	-9	317	311	2	5	9	166	-165
1	4	22	-114	21	4	4	-4	-112	3	7	4	-18	256	-246	5	4	-8	433	-415	2	5	8	419	415
1	4	23	271	267	4	4	-3	411	-421	7	4	-17	373	-364	10	4	-7	225	215	2	5	7	1293	1210
2	4	-26	591	576	4	4	-2	514	510	7	4	-16	239	246	10	4	-6	-124	42	2	5	6	728	-708
2	4	-15	361	375	4	4	-1	197	-113	7	4	-15	577	593	10	4	-5	-117	-87	2	5	5	354	356
2	4	-18	695	685	4	4	0	152	-158	7	4	-14	755	761	10	4	-4	131	133	2	5	4	658	672
2	4	-17	-110	51	4	4	1	-119	15	7	4	-13	504	-495	10	4	-3	400	-435	2	5	3	432	-421
2	4	-16	-115	63	4	4	2	214	215	7	4	-12	511	-511	10	4	-2	179	192	2	5	2	144	140
2	4	-15	813	-804	4	4	3	1303	1308	7	4	-11	132	98	10	4	-1	138	-162	2	5	1	-119	-66
2	4	-14	137	142	4	4	4	115	-1326	7	4	-10	115	-113	10	4	0	534	534	2	5	0	354	356
2	4	-13	522	-365	4	4	5	371	369	7	4	-9	763	-763	10	4	1	418	411	2	5	1	618	-619
2	4	-12	564	-587	4	4	6	421	-424	7	4	-8	318	-305	10	4	2	651	-559	2	5	2	351	347
2	4	-11	534	554	4	4	7	215	-311	7	4	-7	553	-556	10	4	3	291	253	2	5	3	92	903
2	4	-10	354	462	4	4	8	217	205	7	4	-6	-115	-96	10	4	4	149	-142	2	5	4	467	-451
2	4	-9	562	575	4	4	9	552	-936	7	4	-5	462	467	10	4	5	156	-97	2	5	5	-111	-91
2	4	-8	724	-702	4	4	10	692	-657	7	4	-4	746	763	10	4	6	-133	145	2	5	6	435	-411
2	4	-7	-115	57	4	4	11	-113	-55	7	4	-3	573	574	10	4	-5	-127	77	2	5	7	935	-947
2	4	-6	812	-861	4	4	12	446	431	7	4	-2	917	-912	10	4	-4	-121	-43	2	5	8	191	206
2	4	-5	1540	-1563	4	4	13	735	741	7	4	-1	-121	-111	10	4	-3	457	-473	2	5	9	193	183
2	4	-4	811	-813	4	4	14	-112	-84	7	4	0	209	-130	10	4	-2	709	-693	2	5	10	338	359
2	4	-3	564	-536	4	4	15	241	245	7	4	1	322	-352	10	4	-1	475	-472	2	5	11	331	-367
2	4	-2	518	-524	4	4	16	353	-373	7	4	2	447	454	10	4	0	373	-375	2	5	12	473	474
2	4	-1	595	564	4	4	17	277	-267	7	4	3	654	-535	10	4	1	655	553	2	5	13	198	623
2	4	0	716	715	4	4	18	423	424	7	4	4	180	-119	10	4	2	638	559	2	5	14	152	-144
2	4	1	164	163	4	4	19	252	-141	7	4	5	142	-141	10	4	3	475	478	2	5	15	439	-469
2	4	2	222	230	5	4	-21	523	-534	7	4	6	765	774	10	4	4	-127	-69	2	5	16	598	-599
2	4	3	393	-252	5	4	-20	-117	126	7	4	7	622	618	10	4	5	-125	-103	2	5	17	529	-525
2	4	4	947	-920	5	4	-19	-103	20	7	4	8	138	-122	10	4	6	435	-654	3	5	19	344	358
2	4	5	565	-567	5	4	-18	171	-192	7	4	9	155	141	10	4	7	341	-427	3	5	10	329	363
2	4	6	164	163	5	4	-17	367	-367	7	4	10	455	-455	10	4	8	445	-442	3	5	11	412	425
2	4	7	510	-511	5	4	-16	473	-485	7	4	11	271	-258	10	4	9	413	-421	3	5	12	-119	-36
2	4	8	137	158	5	4	-15	227	230	7	4	12	251	251	10	4	10	329	-329	3	5	13	367	368
2	4	9	575	546	5	4	-14	714	-727	7	4	13	189	-213	10	4	11	338	336	3	5	14	334	-344
2	4	10	223	-225	5	4	-13	-135	-146	7	4	14	798	-240	10	4	12	159	232	3	5	15	136	-148
2	4	11	116	162	5	4	-12	118	-118	8	4	-1	612	-612	10	4	13	-127	27	3	5	16	144	-144
2	4	12	262	215	5	4	-11	769	-761	8	4	-19	148	156	12	4	-7	389	349	3	5	11	515	-520
2	4	13	-112	-81	5	4	-10	755	757	8	4	-13	-114	82	12	4	-6	240	-223	3	5	10	-113	113
2	4	14	514	-518	5	4	-9	251	-271	8	4	-17	466	-453	12	4	-5	311	321	3	5	9	547	559
2	4	15	642	-646	5	4	-8	462	-457	8	4	-16	848	833	0	5	-1	422	441	3	5	8	745	742
2	4	16	374	374	5	4	-7	1157	1180	8	4	-15	442	-428	0	5	2	711	-733	3	5	7	-192	-171
2	4	17	415	-409	5	4	-6	1159	1351	8	4	-14	259	-238	0	5	3	534	539	3	5	6	-136	73
2	4	18	-112	-2	5	4	-5	263	314	8	4	-13	475	475	0	5	4	457	-467	3	5	5		

Table 3. Continued.

H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)
4	5	-4	213	-235	6	5	-6	242	236	8	5	3	-128	75	1	6	6	760	-752
4	5	-3	374	-356	6	5	-5	230	258	8	5	4	203	-226	1	6	7	298	194
4	5	-2	237	-345	6	5	-4	574	248	8	5	5	419	-414	1	6	8	-115	93
4	5	-1	-121	80	6	5	-3	138	-130	8	5	6	-129	62	1	6	9	-119	-101
4	5	0	167	-205	6	5	-2	352	324	8	5	7	-121	92	1	6	10	-114	-137
4	5	1	618	-624	6	5	-1	534	539	9	5	-13	434	-421	1	6	11	135	-133
4	5	2	223	254	6	5	0	515	-526	9	5	-12	-124	63	1	6	12	165	152
4	5	3	262	-374	6	5	1	505	533	9	5	-11	351	345	2	6	-14	-121	-94
4	5	4	740	743	6	5	2	505	-458	9	5	-10	292	288	2	6	-13	130	-94
4	5	5	564	524	6	5	3	656	-685	9	5	-9	-121	75	2	6	-12	738	-749
4	5	6	-116	-24	6	5	4	294	309	9	5	-8	161	166	2	6	-11	121	-135
4	5	7	-116	17	6	5	5	-127	42	9	5	-7	415	346	2	6	-10	308	403
4	5	8	582	-551	6	5	6	223	250	9	5	-6	514	-516	2	6	-9	215	-232
4	5	9	277	274	6	5	7	144	-143	9	5	-5	-130	-79	2	6	-8	218	234
4	5	10	-122	-154	6	5	8	-121	-95	9	5	-4	527	-518	2	6	-7	-111	229
4	5	11	-121	-112	6	5	9	-137	121	9	5	-3	723	-715	2	6	-6	254	225
4	5	12	254	255	6	5	10	237	-250	9	5	-2	223	222	2	6	-5	412	324
4	5	13	572	-554	6	5	11	441	416	9	5	-1	361	279	2	6	-4	291	274
4	5	14	246	241	7	5	-17	-114	12	9	5	0	348	357	2	6	-3	137	115
5	5	-15	173	146	7	5	-16	153	-162	9	5	1	160	-105	2	6	-2	1010	-1099
5	5	-16	130	-115	7	5	-15	553	-559	9	5	2	158	195	2	6	-1	-127	57
5	5	-17	-112	-64	7	5	-14	927	543	9	5	3	171	158	2	6	0	-158	6
5	5	-18	335	332	7	5	-13	151	215	10	5	-9	-144	-44	2	6	1	-126	59
5	5	-19	147	123	7	5	-12	426	452	10	5	-8	139	130	2	6	2	714	693
5	5	-20	284	283	7	5	-11	451	462	10	5	-7	266	212	2	6	3	-120	-63
5	5	-21	645	644	7	5	-10	214	-215	10	5	-6	193	-205	2	6	4	-117	31
5	5	-22	254	-224	7	5	-9	225	-235	10	5	-5	173	-184	2	6	5	-119	69
5	5	-23	711	-658	7	5	-8	176	-165	10	5	-4	267	-210	2	6	6	244	289
5	5	-24	491	-472	7	5	-7	356	301	10	5	-3	178	200	2	6	7	-115	41
5	5	-25	-115	-115	7	5	-6	-121	-72	10	5	-2	135	-133	2	6	8	939	-937
5	5	-26	-375	355	7	5	-5	417	-428	10	5	0	816	824	2	6	9	-122	-194
5	5	-27	-112	-35	7	5	-4	657	252	10	5	1	-120	138	2	6	10	-128	-125
5	5	-28	64	645	7	5	-3	-131	100	10	5	2	-112	-55	2	6	11	118	-233
5	5	-29	278	-269	7	5	-2	463	385	10	5	3	351	362	2	6	12	177	-153
5	5	-30	277	339	7	5	-1	870	601	10	5	4	316	-304	2	6	13	-113	-15
5	5	-31	146	143	7	5	0	-137	-53	10	5	5	152	115	2	6	14	220	-211
5	5	-32	476	-511	7	5	1	330	-301	10	5	6	342	-371	2	6	15	219	-225
5	5	-33	281	-232	7	5	2	539	-567	10	5	7	150	-157	2	6	16	459	-441
5	5	-34	261	-247	7	5	3	-145	-56	10	5	8	170	-212	2	6	17	152	137
5	5	-35	1258	-1253	7	5	4	231	-280	10	5	9	126	150	2	6	18	212	202
5	5	-36	254	237	7	5	5	238	-236	10	5	10	766	758	2	6	19	279	319
5	5	-37	442	437	7	5	6	229	232	10	5	11	-119	-5	2	6	20	792	717
5	5	-38	-129	136	7	5	7	394	-258	10	5	12	-124	141	2	6	21	-110	5
5	5	-39	333	-234	7	5	8	411	425	10	5	13	-117	-45	2	6	22	297	-299
5	5	-40	420	412	7	5	9	349	354	11	6	-13	173	-174	2	6	23	135	134
5	5	-41	64	64	8	5	-16	-122	-94	11	6	-12	211	240	2	6	24	-134	-113
5	5	-42	240	-257	8	5	-15	154	144	11	6	-11	-114	3	2	6	25	-112	34
5	5	-43	-125	88	8	5	-14	-142	-45	11	6	-10	234	-215	2	6	26	254	-261
5	5	-44	578	-560	8	5	-13	136	-125	11	6	-9	-118	44	2	6	27	-131	-92
5	5	-45	511	-517	8	5	-12	470	466	11	6	-8	739	732	2	6	28	294	-173
5	5	-46	-134	-21	8	5	-11	252	-248	11	6	-7	235	-225	2	6	29	-123	51
5	5	-47	256	278	8	5	-10	505	523	11	6	-6	-105	15	2	6	30	723	703
6	5	-16	155	-153	8	5	-9	812	905	11	6	-5	153	-145	2	6	31	166	-155
6	5	-17	374	-383	8	5	-8	454	-444	11	6	-4	945	-905	2	6	32	-118	33
6	5	-18	134	131	8	5	-7	137	-135	11	6	-3	-114	37	2	6	33	160	-155
6	5	-19	169	134	8	5	-6	353	-404	11	6	-2	474	285	2	6	34	214	-197
6	5	-20	246	234	8	5	-5	-131	106	11	6	-1	-123	81	2	6	35	-122	-33
6	5	-21	217	-204	8	5	-4	-123	48	11	6	0	-134	69	2	6	36	10	275
6	5	-22	-114	76	8	5	-3	-125	-51	11	6	1	246	221	2	6	37	-121	123
6	5	-23	252	214	8	5	-2	205	210	11	6	2	395	332	2	6	38	-115	-12
6	5	-24	324	-324	8	5	-1	708	-710	11	6	3	-113	17	2	6	39	151	189
6	5	-25	-111	-12	8	5	0	202	190	11	6	4	227	207	2	6	40	-10	657
6	5	-26	212	-158	8	5	1	712	706	11	6	5	-111	14	2	6	41	-119	81
6	5	-27	670	-666	8	5	2	315	-285										

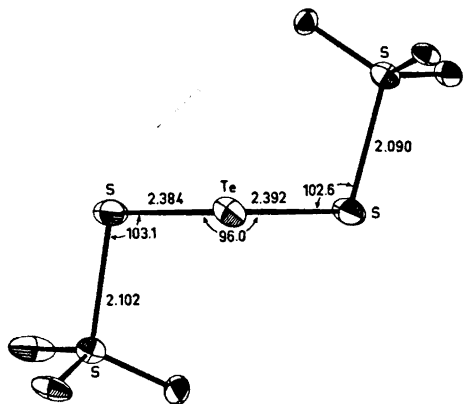


Fig. 1. The *trans* form of the telluropentathionate ion in $\text{BaTe}(\text{S}_2\text{O}_5)_2 \cdot 3\text{H}_2\text{O}$ as seen along the *b* axis. The ellipsoids represent 50% probability; the thermal parameters of the oxygen atoms are halved in size relative to those of the tellurium and sulphur atoms.

Table 4. Dimensions of the telluropentathionate ion. Standard deviations, including estimated uncertainties in unit cell dimensions, are given in parentheses.

Bond lengths and angles	
S(1)–S(2) = 2.1020(16) Å	S(4)–S(5) = 2.0898(16) Å
S(2)–Te = 2.3835(15)	Te–S(4) = 2.3924(16)
∠S(1)–S(2)–Te = 103.08(5)°	∠Te–S(4)–S(5) = 102.59(5)°
∠S(2)–Te–S(4) = 96.02(4)°	
S(1)–O(1) = 1.427(4) Å	S(5)–O(4) = 1.451(3) Å
S(1)–O(2) = 1.428(4)	S(5)–O(5) = 1.456(4)
S(1)–O(3) = 1.432(4)	S(5)–O(6) = 1.447(4)
∠S(2)–S(1)–O(1) = 107.2(2)°	∠S(4)–S(5)–O(4) = 106.7(2)°
∠S(2)–S(1)–O(2) = 102.2(2)°	∠S(4)–S(5)–O(5) = 102.0(2)°
∠S(2)–S(1)–O(3) = 108.2(2)°	∠S(4)–S(5)–O(6) = 109.1(2)°
∠O(1)–S(1)–O(2) = 114.3(3)°	∠O(4)–S(5)–O(4) = 113.5(2)°
∠O(1)–S(1)–O(3) = 111.1(3)°	∠O(4)–S(5)–O(6) = 112.5(2)°
∠O(2)–S(1)–O(3) = 113.1(3)°	∠O(5)–S(6)–O(6) = 112.1(2)°
Dihedral angles	
S(1)S(2)Te/S(2)TeS(4) = 87.7°	S(2)TeS(4)/TeS(4)S(5) = 98.7°
TeS(2)S(1)/S(2)S(1)O(1) = 70.0°	TeS(4)S(5)/S(4)S(5)O(4) = 68.8°
TeS(2)S(1)/S(2)S(1)O(2) = 169.5°	TeS(4)S(5)/S(4)S(5)O(5) = 171.9°
TeS(2)S(1)/S(2)S(1)O(3) = 50.0°	TeS(4)S(5)/S(4)S(5)O(6) = 53.1°
S(2)S(1)O(1)/S(2)S(1)O(2) = 120.5°	S(4)S(5)O(4)/S(4)S(5)O(5) = 119.3°
S(2)S(1)O(1)/S(2)S(1)O(3) = 119.9°	S(4)S(5)O(4)/S(4)S(5)O(6) = 121.9°
S(2)S(1)O(2)/S(2)S(1)O(3) = 119.6°	S(4)S(5)O(5)/S(4)S(5)O(6) = 118.8°
Non-bonded distances	
S(1)–Te = 3.5168(17) Å	Te–S(5) = 3.5028(18) Å
S(1)–S(4) = 4.3747(22)	S(2)–S(5) = 4.5652(23)
S(2)–S(4) = 3.5497(21)	S(1)–S(5) = 5.9914(27)

on film data⁸ are about 1.5 times the standard deviations estimated for the film data values.

The non-planar S–S–Te–S–S chain in the *trans* form possesses the symmetry of a twofold axis, when the two halves of the chain are identical. This is the case in the crystals of tellurium dibenzenethiosulphonate,¹⁷ where a twofold axis is crystallographically required.

In the present salt the crystal symmetry does not impose a twofold axis, but the dimensions are, except for the dihedral angles, nearly the same in the two halves of the telluropentathionate ion. The S–S and S–Te bond lengths, and the S–S–Te bond angles differ by only 0.0122 Å, 0.0089 Å, and 0.49°, respectively, and thus correspond nearly to the symmetry of a twofold axis. The dihedral angles between the S–S–Te and S–Te–S planes, however, differ by 11.0°. This difference in the degrees of rotation about the S–Te bonds is also seen from the non-bonded S(1)–S(4) and S(2)–S(5) distances which differ by 0.1905 Å. The dimensions of the two distorted tetrahedrally shaped thiosulphate groups are also nearly the same, and as seen from the

dihedral angles of Table 4, there is only a small difference in the degrees of rotation of the sulphonate groups about the S—S bonds.

The oxygen atoms are numbered in the same order as in the description of the structure of barium telluropentathionate dihydrate.⁷ The dihedral angles thus show that the degrees of rotation of the sulphonate groups about the S—S bonds are different by about 21° in the two structures.

Table 5 gives the dimensions of the *trans* form of the S—S—Te—S—S chain in three compounds. The greatest variations are again found in the

Table 5. Bond lengths (Å), bond angles (°) and dihedral angles (°) of the S—S—Te—S—S chain in the *trans* form. Standard deviations are given in parentheses.

	BaTe(S ₂ O ₃) ₂ ·3H ₂ O	Rb ₂ Te(S ₂ O ₃) ₂ ·1½H ₂ O ¹⁰	Te(S ₂ O ₂ C ₆ H ₅) ₂ ¹⁷
S—S	2.102(2), 2.090(2)	2.116(11), 2.126(9)	2.080(2)
S—Te	2.384(2), 2.392(2)	2.364(9), 2.370(7)	2.380(2)
∠S—S—Te	103.08(5), 102.59(5)	103.3(4), 102.8(3)	103.46(7)
∠S—Te—S	96.02(4)	100.1(3)	97.71(6)
∠SSTe/STeS	87.7, 98.7	77.7, 89.2	97.4

SSTe/STeS diredral angles. In each of the two salts, the two angles within a chain differ by about 11°, and the average values in the two salts differ by about 10°. The larger dihedral angles in tellurium dibenzenethiosulphonate might be due to the space requirement of the benzenesulphonate groups. The variations in the two salts, however, show that the dihedral angles are also influenced by the cation-oxygen coordination or by other packing effects. The variations in S—S and S—Te bond lengths, and in S—S—Te bond angles are small, though significant. The S—Te—S bond angles differ from 96.02(4)° in the barium salt to 100.1(3)° in the rubidium salt.

Among the barium salts of penta-, selenopenta- and telluropentathionic acids, BaTe(S₂O₃)₂·3H₂O is the only example where the anion occurs in the *trans* form. In the crystals where the *cis* form is found, each barium ion is in close contact with nine oxygen atoms, of which three are from the water molecules and the remaining six are from the anions. Two of the latter contacts are from different sulphonate groups of the same anion. All the oxygen atoms, except that of one of the water molecules in BaSe(S₂O₃)₂·3H₂O,⁵ are involved in the Ba—O coordination. The Ba—O distances are in the range 2.75–2.94 Å.^{4,5,7}

In the present structure the barium ions are each surrounded by eight oxygen atoms in the range 2.69–2.85 Å. All the water and sulphonate oxygen atoms, except O(4), take part once in these approaches. In addition there is one Ba—O distance of 3.19 Å, this oxygen atom being from the same sulphonate group as one of those involved in the closer contacts. The individual distances and related angles are listed in Table 6. Fig. 2 is a stereoscopic view as seen normal to the *b* crystal plane, showing the content of two

Table 6. Distances (Å), and angles (°) between directions, from barium ion to oxygen atoms. Standard deviations of the distances and angles are 0.003–0.005 Å and 0.1°, respectively.

			I	II	III	IV	V	VI	VII	VIII
I	H ₂ O(1)	(1+x, 1+y, z)	2.794							
II	H ₂ O(2)	(1+x, y, z)	2.797	71.7						
III	H ₂ O(3)	(x, y, z)	2.838	148.1	131.7					
IV	O(1)	(1+x, y, z)	2.783	68.1	69.2	135.0				
V	O(2)	(1-x, 2-y, z)	2.693	70.5	141.9	85.2	91.9			
VI	O(3)	(1-x, 1-y, z)	2.808	123.1	126.8	64.6	71.7	72.4		
VII	O(5)	(x, y, z)	3.190	131.1	59.5	76.5	91.6	157.4	87.6	
VIII	O(5)	(x, 1+y, z)	2.848	71.2	103.7	81.1	138.9	68.6	129.4	120.4
IX	O(6)	(x, y, z)	2.824	120.3	71.7	63.6	133.8	134.2	116.5	46.7

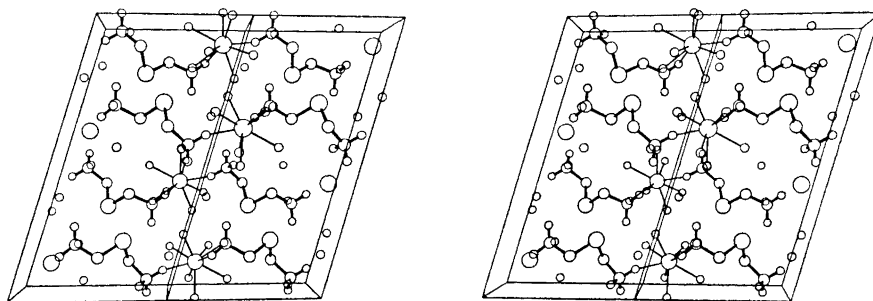


Fig. 2. A stereoscopic view as seen normal to the *b* crystal plane. The Ba–O coordination is indicated by the thin lines.

unit cells along the *a* axis. The Ba–O contacts are indicated by thin lines, whereas the bonds within the telluropentathionate ions are indicated by thick lines. The sulphonate oxygen atoms coordinated to the two barium ions in the middle of the figure are from the four nearest telluropentathionate ions shown, or from ions generated from the latter by addition or subtraction of whole *b* units. The Ba–O coordination thus forms layers parallel to the *c* crystal plane, with thickness half the *c* spacing. The relation between neighbour layers is through twofold screw axes at $z = \frac{1}{4}$ and $\frac{3}{4}$. There are no Ba–O coordination contacts across the interfaces between the layers. This is very similar to the situation found for BaTe(S₂O₃)₂·2H₂O,⁷ and the other barium salts with anions in the *cis* form, and where the crystals show perfect cleavage along the *c* crystal planes.¹⁸ There is no such tendency of cleavage in the present crystals, and this might be due to the closer packing of the layers, to hydrogen bonds or other interionic contacts between the layers.

The most probable hydrogen bonds are; from H₂O(1) to O(1) at (*x*, *y*, *z*) and (*x*, *y* – 1, *z*) with distances 2.993(6) Å and 3.126(6) Å and angle 117.9(2)°, from H₂O(2) to H₂O(1) at (*x*, *y*, *z*) and O(4) at ($1 - x, \frac{1}{2} + y, \frac{1}{2} - z$) with distances

2.917(6) Å and 2.864(5) Å and angle 99.3(2)°, and from H₂O(3) to O(6) at (x,y,z) and O(3) at (1-x,1-y,z) with distances 2.984(5) Å and 3.018(7) Å and angle 105.9(2)°. Only one of these assumed hydrogen bonds, H₂O(2)-O(4), is across the interface between barium-oxygen coordination layers.

The closest interionic Te...Te approaches are between the tellurium atom at (x,y,z) and those at (1-x,½+y,½-z) and (1-x,-½+y,½-z). The distances are 3.833(1) Å and the Te...Te...Te angle is 86.31(3)°. The oxygen atom O(4), not involved in the Ba-O coordination, at (1-x,½+y,½-z), is at a distance of 2.926(4) Å from tellurium. This oxygen atom and the tellurium at (1-x,-½+y,½-z) are located only 0.11 Å and 0.40 Å, respectively, out of the S(2)-Te-S(4) plane. The angles between directions from tellurium to its four neighbours in this plane are: ∠S(2)-Te-S(4)=96.02(4)°, ∠S(4)-Te...Te=69.49(4)°, ∠Te...Te...O(4)=113.55(6)°, ∠O(4)...Te-S(2)=80.63(7)°, ∠S(2)-Te...Te=164.25(4)°, and ∠S(4)-Te...O(4)=175.86(8)°. There are thus two approximately linear systems, S(2)-Te...Te and S(4)-Te...O(4), at nearly right angles. The Te...Te and Te...O approaches are 0.29 Å and 0.65 Å, respectively, shorter than the sum of the van der Waals radii of the atoms involved.¹⁹ The Te...Te and Te...O contacts are across the interfaces between the layers mentioned earlier.

REFERENCES

1. Foss, O. and Zachariasen, H. *Acta Chem. Scand.* **8** (1954) 473.
2. Foss, O. and Tjomsland, O. *Acta Chem. Scand.* **10** (1956) 288.
3. Foss, O. and Tjomsland, O. *Acta Chem. Scand.* **8** (1954) 1701.
4. Marøy, K. *Acta Chem. Scand.* **26** (1972) 36.
5. Marøy, K. *Acta Chem. Scand.* **26** (1972) 45.
6. Foss, O. and Tjomsland, O. *Acta Chem. Scand.* **12** (1958) 52.
7. Marøy, K. *Acta Chem. Scand.* **27** (1973) 1695.
8. Gjerrestad, K. and Marøy, K. *Acta Chem. Scand.* **24** (1970) 3402.
9. Foss, O. and Larssen, P. A. *Acta Chem. Scand.* **8** (1954) 1042.
10. Marøy, K. *Acta Chem. Scand.* **25** (1971) 2557.
11. Foss, O. and Tjomsland, O. *Acta Chem. Scand.* **10** (1956) 416.
12. Coppens, P., Leiserowitz, L. and Rabinovich, D. *Acta Cryst.* **18** (1965) 1035.
13. Zachariasen, W. H. *Acta Cryst.* **16** (1963) 1139.
14. *International Tables for X-Ray Crystallography*, Kynoch Press, Birmingham 1962, Vol. III.
15. Thomas, L. H. and Umeda, K. *J. Phys. Chem.* **26** (1957) 293.
16. Cromer, D. T. *Acta Cryst.* **18** (1965) 17.
17. Åse, K. *Acta Chem. Scand.* **25** (1971) 838.
18. Foss, O. *Advan. Inorg. Chem. Radiochem.* **2** (1960) 237.
19. Bondi, A. J. *J. Phys. Chem.* **68** (1964) 441.

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